

Design, Construction, and Testing of Particle Sources

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Supervised by Clark Chen and

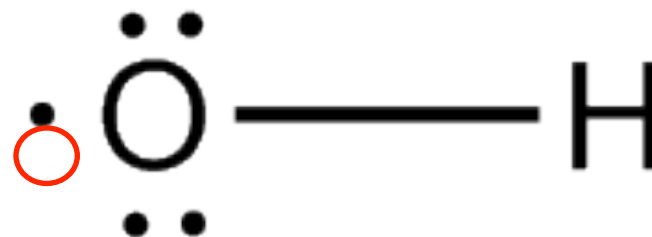
Prof. Koel

Outline

- Background
- Introduction
- Atomic Hydrogen Dosers
 - Tungsten Tube
 - Platinum Tube
 - Reference Data from PPPL
- Methyl Radical Doser
 - One Heated Zone
 - Two Heated Zone
- Future Work

Background

- To study plasma/surface interactions, require sources of controlled energetic particles
- Particle sources – apparatuses that split molecules into radicals (have “dangling” covalent bond)
 - Radicals are highly reactive



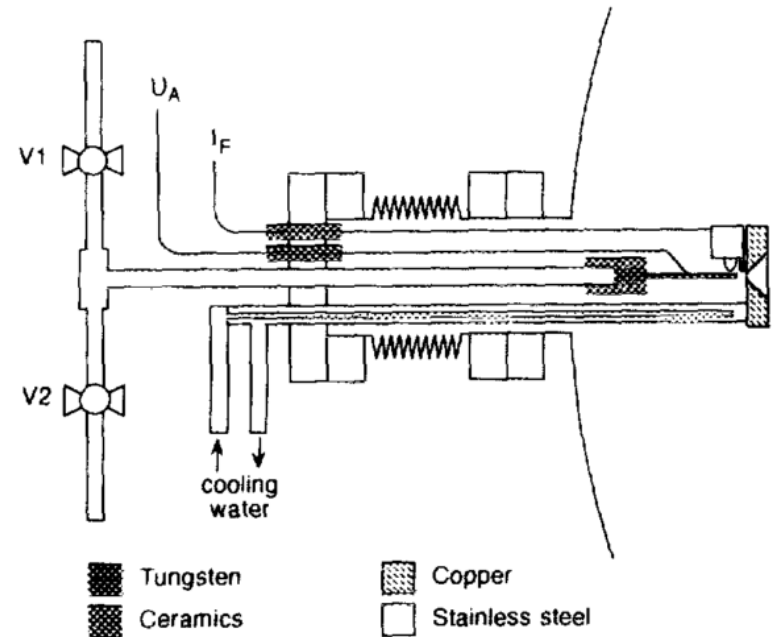
Hydroxyl Radical – red circle highlights missing electron

Introduction

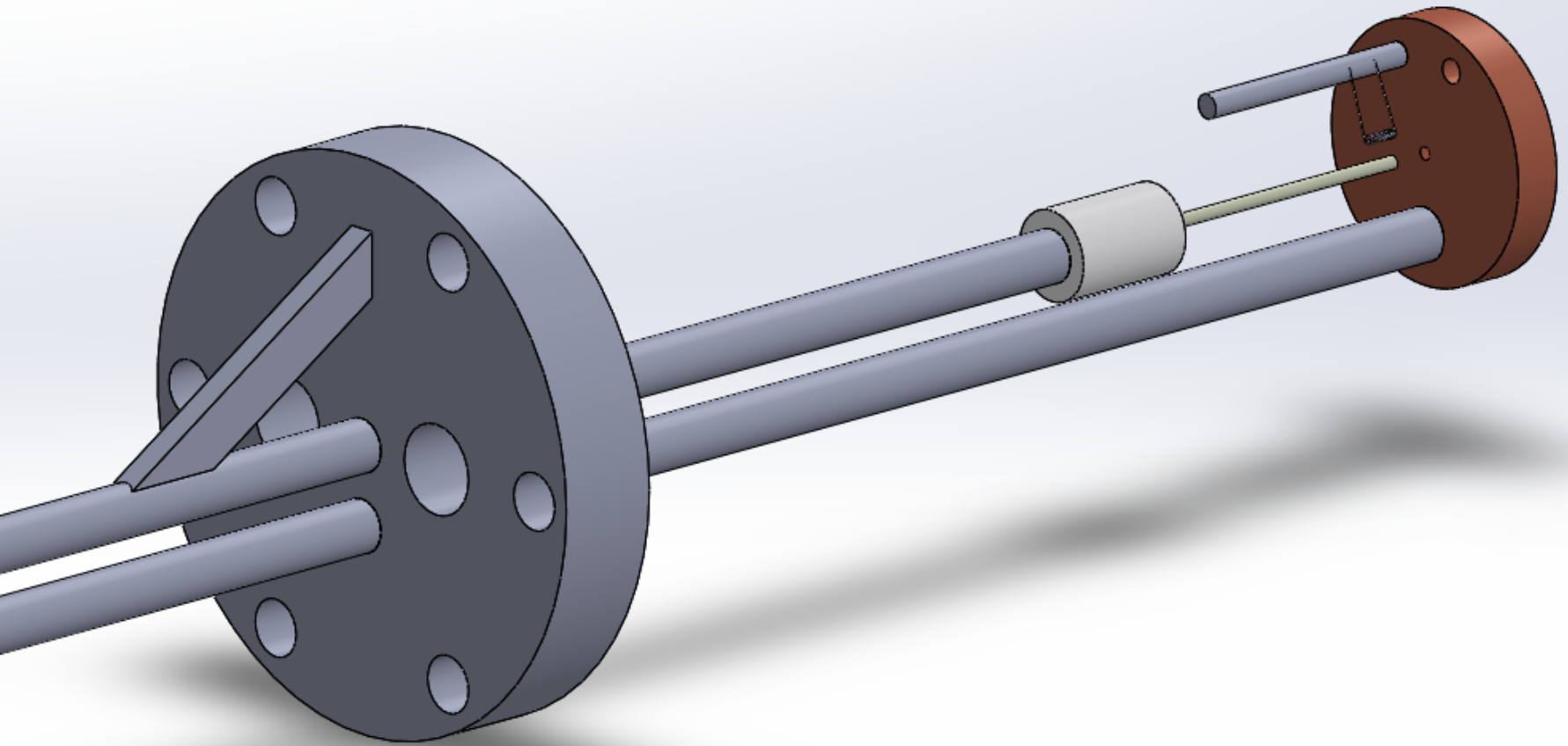
- To compare efficiencies of different types, look at degree of dissociations
- Worked on:
 - Atomic Hydrogen dosers
 - Methyl Radical dosers
 - Compared physical and material constraints of each design

Atomic H Dosers (Tungsten Tube)

- Designed by Bischler and Bertel
- Tip of tungsten tube is heated via electron bombardment by nearby tungsten filament
- Copper shield acts as both thermal protection and collimator



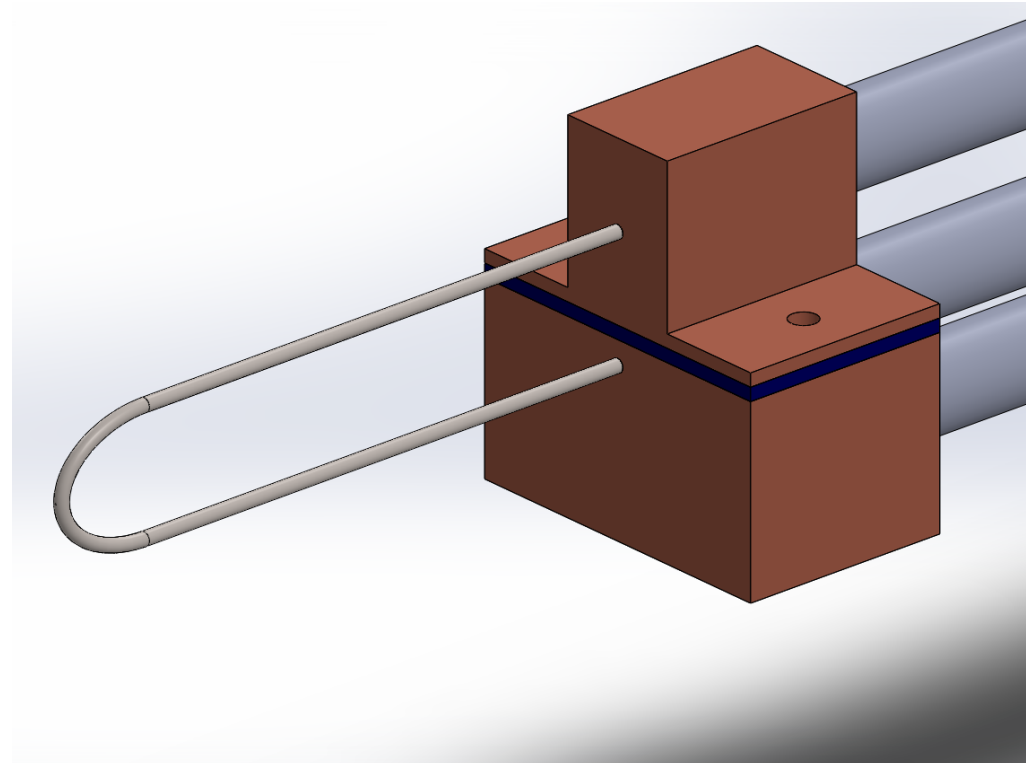
Cross section schematic view of doser
U. Bischler and E. Bertel, J. Vac. Sci. Technol. **11**, 458
(1993)



3D view of doser based on B&B's design

Atomic H Dosers (Platinum Tube)

- Based on design by Engel and Rieder
- Platinum tube is resistively heated
- Copper blocks electrically but not thermally isolated, allowing for cooling
- Hole in platinum tube shoots out H atoms



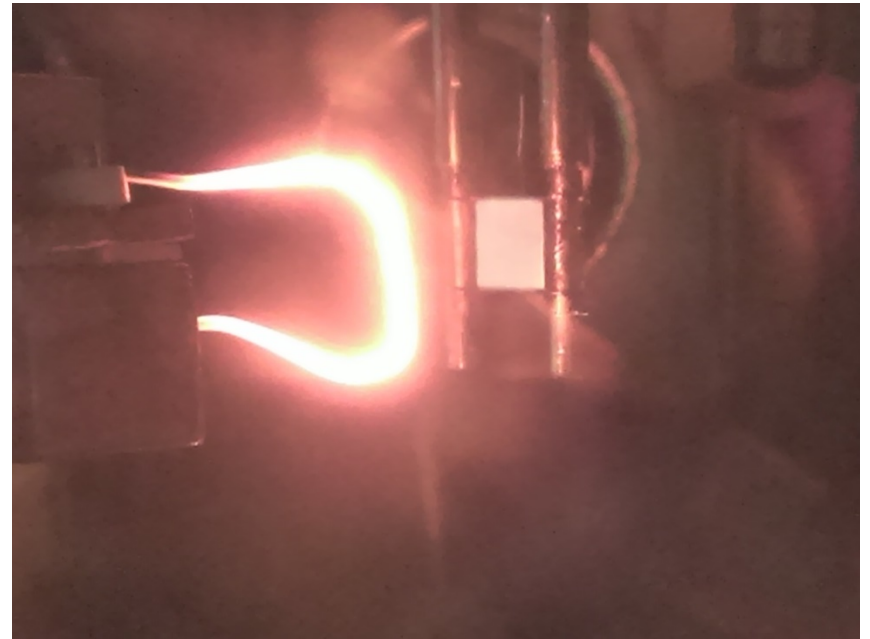
3D view of important components of doser

Why use Pt. Tube?

- Easier to make
- Less expensive (tungsten tube vs. platinum tube)
- Platinum inert to certain corrosive gases that would destroy tungsten tube (M.R. Voss, H. Busse, and B.E. Koel, Surf. Sci, **414**, 330 (1998))

Reference Data from PPPL

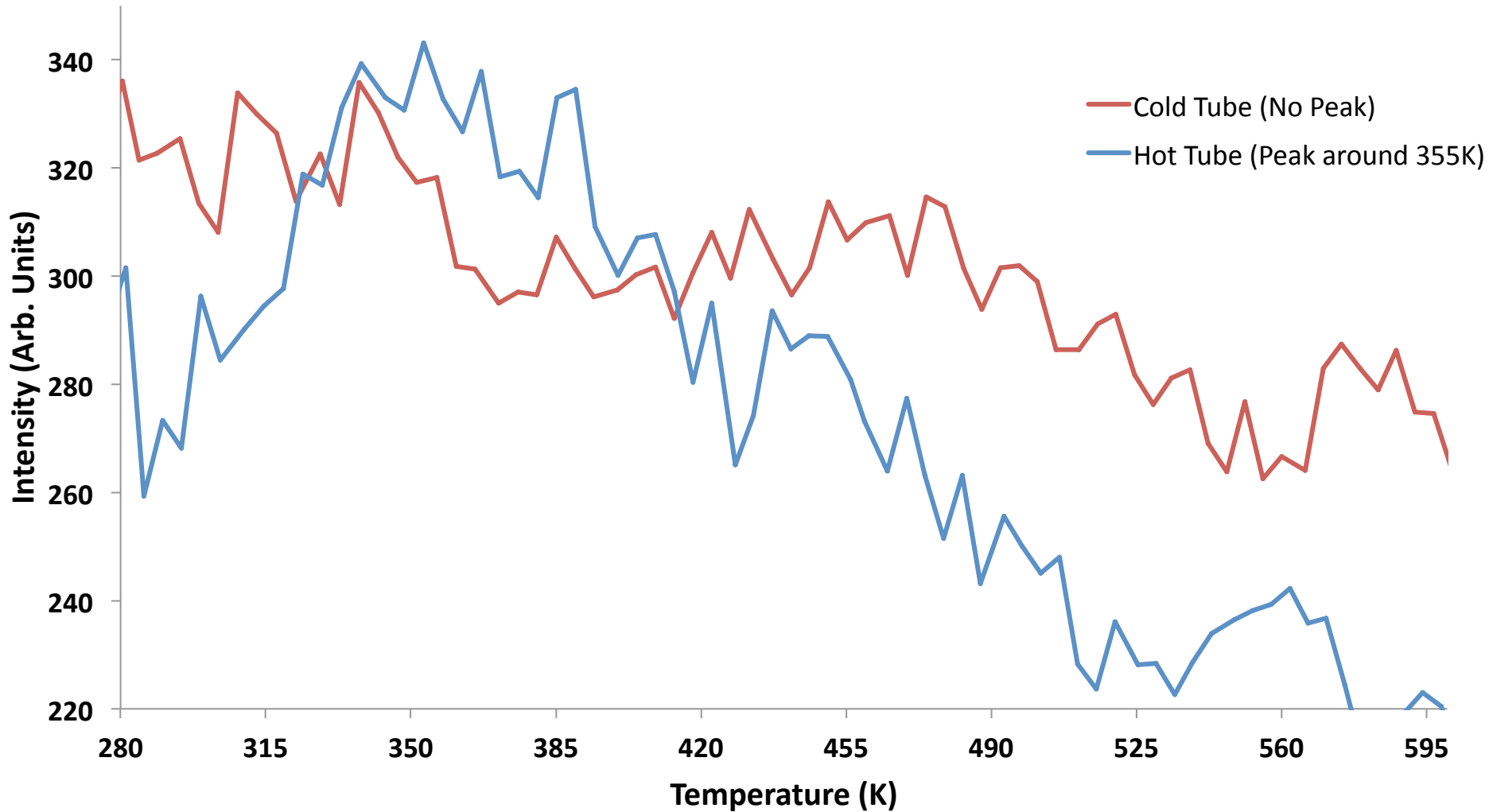
- Substrate is TZM
 - Titanium zirconium alloy
 - Not ideal
- Dissociating D_2
- Control
 - Cold tube, initial temp at 265.5K, $2e-7$ torr for 50s = 10 LM
- Experimental
 - Hot tube (shown), initial temp at 267.0K, 10LM



PPPL doser in operation

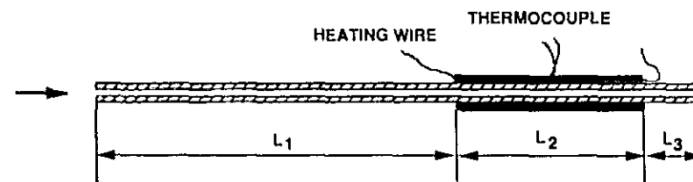
D₂ (mass of 3.725) on TZM

10K/s

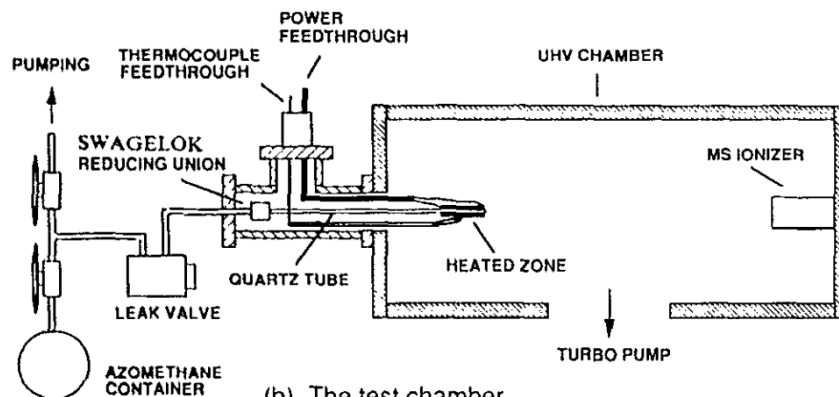


Methyl Radical Doser (One Heated Zone)

- Based on Peng, et. al. design
- Tantalum wire wrapped around quartz tube resistively heated, covered with alumina-based cement
- Max temp of 1300K with 10A running through
- Have one in lab
 - Working condition - ??



(a) The tubular reactor



(b) The test chamber

Schematic of doser

X. D. Peng, R. Viswanathan, G. Smudde, and P. Stair, Rev. Sci. Instrum. **63**, 3930 (1992)

Methyl Radical Doser (Two Heated Zones)

- Based on Dickens and Stair's design
- As before, but two heated zones
- Can independently heat the two zones
 - More control over degree of dissociation and exiting temperature of methyl radicals

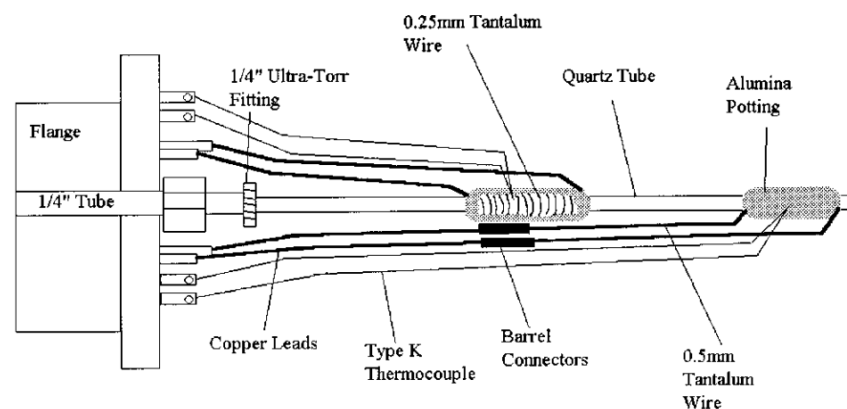
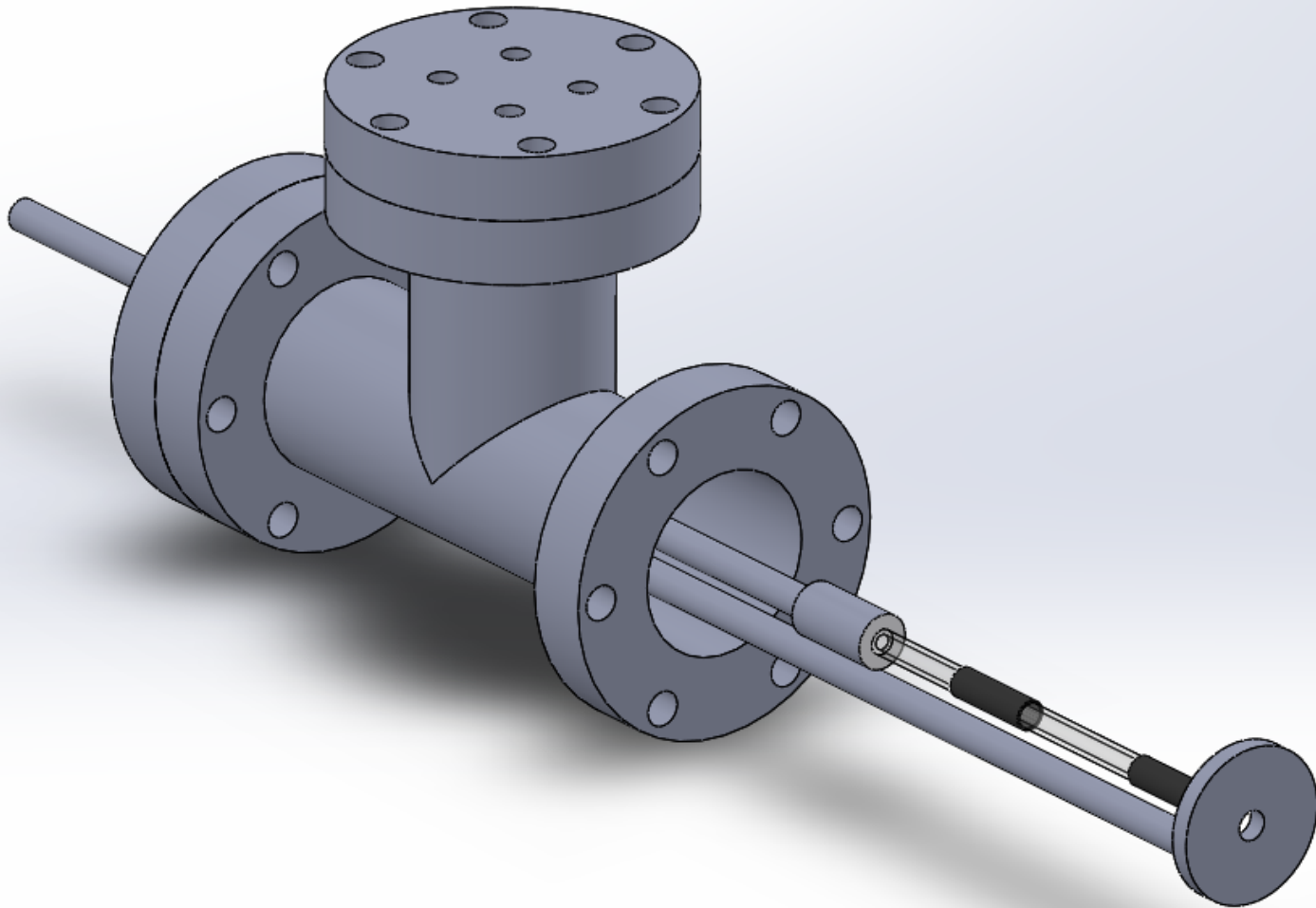
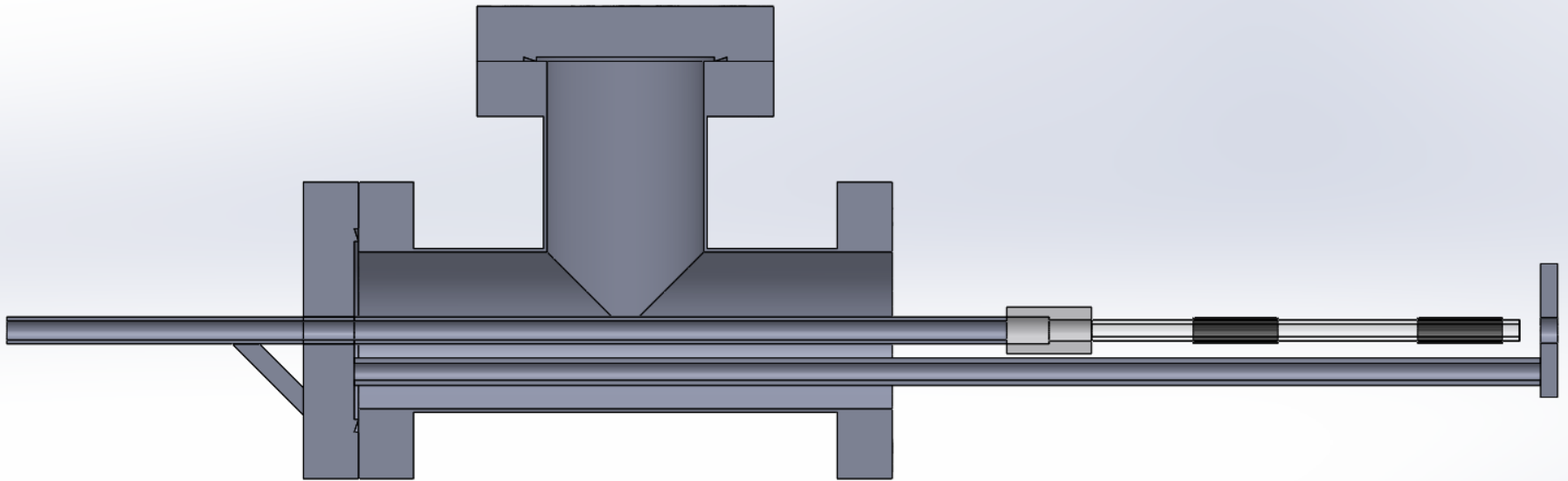


Figure 7. Schematic of doser
K.A. Dickens and P. Stair, *Langmuir*, **14**,
1444 (1998)



Proposed Methyl Radical Doser with Two Heated Zones Design
(Isometric View)



Proposed Methyl Radical Doser with Two Heated Zones Design
(Side cross-section view)

Future Work and Experiments

- Current Status of Dosers –
 - Tungsten tube H doser – repaired and ready to be tested
 - Platinum tube H doser – ready to be assembled via silver soldering (two more have been repaired, one need sapphire plate, and both need platinum tubes)
 - Methyl radical doser – In design stages, need to finalize design and order flange

References

- Bischler, U. (1993). Simple source of atomic hydrogen for ultrahigh vacuum applications. *Journal Of Vacuum Science & Technology A: Vacuum, Surfaces, And Films*, 11(2), 458. Retrieved from <http://scitation.aip.org/content/avs/journal/jvsta/11/2/10.1116/1.578754>
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- Peng, X., Viswanathan, R., Smudde, G., & Stair, P. (1992). A methyl free radical source for use in surface studies. *Review Of Scientific Instruments*, 63(8), 3930. Retrieved from <http://scitation.aip.org/content/aip/journal/rsi/63/8/10.1063/1.1143240>
- Voss, M. (1998). *Adsorption and Reaction of Small Molecules and Radicals on Metal Alloys and Molecular Solids* (PhD). Princeton University.